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Combined Gas Law Worksheet

Show your work! The numerical answers (before counting for significant digits) are at the bottom
→ they are just for you to check yourself! I still need to see all your work!

1. I have an unknown volume of gas at a pressure of 0.5 atm and a temperature of 325 K. If I raise the pressure to 1.2 atm, decrease the temperature to 320 K, and measure the final volume to be 48 liters, what was the initial volume of the gas?
2. A toy balloon has an internal pressure of 1.05 atm and a volume of 5.0 L. If the temperature where the balloon is released is 20° C, what will happen to the volume when the balloon rises to an altitude where the pressure is 0.65 atm and the temperature is -15° C?
3. A small research submarine with a volume of 1.2×10^3 L has an internal pressure of 1.0 atm and an internal temperature of 15° C. If the submarine descends to a depth where the pressure is 150 atm and the temperature is 3° C, what will the volume of the gas inside be if the hull of the submarine breaks?
4. A gas that has a volume of 28 liters, a temperature of 45° C, and an unknown pressure has its volume increased to 34 liters and its temperature decreased to 35° C. If I measure the pressure after the change to be 2.0 atm, what was the original pressure of the gas?

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